Limiting Reactant and Excess Reactant Name \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Practice

1. The reaction between sodium and iron(III) oxide is one in a series of reactions that inflates an airbag: 6 Na + Fe2O3 → 3 Na2O + 2 Fe If 100.0 g of Na and 100.0 g of Fe2O3 are used in this reaction, determine the following:
2. Limiting reactant =
3. Reactant in excess =
4. Mass of solid iron produced =
5. Mass of excess reactant that remains after the reaction is complete =
6. Photosynthesis reactions in green plants use carbon dioxide and water to produce glucose (c6H12O6) and oxygen. A plant has 88.0 g of carbon dioxide and 64.0 g of water available for photosynthesis.
7. Write the balanced equation for the reaction:
8. Limiting reactant =
9. Excess reactant =
10. Mass of glucose produced =
11. Mass of excess reactant that remains after the reaction is complete =

Answer Key

