Matter and Chemical Change Lab Name \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

1. Video clip 1 – write down one evidence that a chemical change occurred

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1. Video clip 2 – write down one evidence that a chemical change occurred

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1. From the counter, grab a zip lock bag, a clear pill container, and a piece of steel wool.
   1. Carefully fill pill container about ½ full with hydrogen peroxide on counter
   2. Place the pill container carefully inside the bag
   3. Place the small piece of steel wool in the bag also without letting it touch the hydrogen peroxide in the pill bottle
   4. Squeeze out as much air from the bag as possible and then carefully zip up the bag
   5. Tip over the pill bottle of hydrogen peroxide so it comes in contact with the steel wool

Write down two evidences that a chemical change occurred

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1. Identify each side of the equation as the products or the reactants
2. Hydrogen + Oxygen Water

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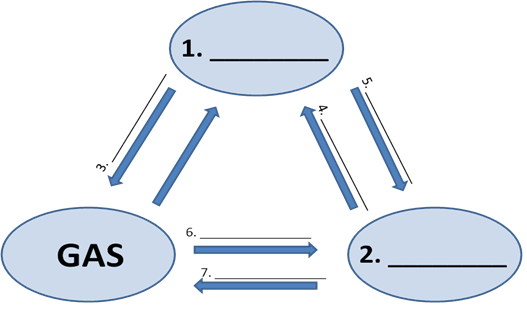
1. Mercury Oxide Liquid Mercury + Oxygen

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1. Sodium Chloride Sodium + Chlorine

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5. Fill out the table with the correct phase states and processes



1. Consider the following decomposition reaction:

2 H2O2 2 H2O + O2

If 72 grams of water and 64 grams of oxygen are produced, what mass of H2O2 decomposed?

1. Consider the following chemical reaction:

2 NaCl + Ca(OH)2 CaCl2 +2 NaOH

If the mass of NaCl reacted is 191 grams and calcium hydroxide 74 grams and 80 grams of sodium hydroxide is produced, what mass of calcium chloride is produced?